

What's in *Your* Water Bottle?

Case Study #1

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References and Further Reading:

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Bottled Water Brand: _____

Source: _____

Treatment: _____

	Conc (mg/L)	Conc (mM)	Conc (meq/L)	
Cations				
Anions				

$$\text{Charge Balance} = (\Sigma -\text{'ve} - \Sigma +\text{'ve}) / (\Sigma -\text{'ve} + \Sigma +\text{'ve}) * 100\%$$

$$\text{Ionic Strength (M)} = 0.5 \Sigma c_i Z_i^2, \text{ where } c_i = \text{molar conc and } Z_i = \text{charge}$$

$$\text{Total Hardness (ppm CaCO}_3) = \Sigma [M^{2+}] \times 100,000 \text{ mg/mol}$$

$$\text{Total Alkalinity (ppm CaCO}_3) = \Sigma [\text{bases}] \times 50,000 \text{ mg/equiv}$$